

# Naming Ionic Compounds Practice Worksheet

Name the following ionic compounds:

1)  $\text{NH}_4\text{Cl}$  Ammonium chloride

2)  $\text{Fe}(\text{NO}_3)_3$  Iron (III) Nitrate

3)  $\text{TiBr}_3$  Titanium (III) Bromide

4)  $\text{Cu}_3\text{P}$  Copper (I) Phosphide

5)  $\text{SnSe}_2$  Tin(IV) Selenide

6)  $\text{GaAs}$  Gallium (III) Arsenide

7)  $\text{Pb}(\text{SO}_4)_2$  Lead(IV) Sulfate

8)  $\text{Be}(\text{HCO}_3)_2$  Beryllium Hydrogen Carbonate

9)  $\text{Mn}_2(\text{SO}_3)_3$  Manganese(III) Sulfite

10)  $\text{Al}(\text{CN})_3$  Aluminum Cyanide

Write the formulas for the following compounds:

11) chromium (VI) phosphate  $\text{Cr}^{+6}\text{PO}_4^{-3} \rightarrow \text{Cr}_2(\text{PO}_4)_3$

12) vanadium (IV) carbonate  $\text{V}^{+4}\text{CO}_3^{-2} \rightarrow \text{V}(\text{CO}_3)_2$

13) tin (II) nitrite  $\text{Sn}^{+2}\text{NO}_2^{-1} \rightarrow \text{Sn}(\text{NO}_2)_2$

14) cobalt (III) oxide  $\text{Co}^{+3}\text{O}^{2-} \rightarrow \text{Co}_2\text{O}_3$

15) titanium (II) acetate  $\text{Ti}^{+2}\text{C}_2\text{H}_3\text{O}_2^{-1} \rightarrow \text{Ti}(\text{C}_2\text{H}_3\text{O}_2)_2$

16) vanadium (V) sulfide  $\text{V}^{+5}\text{S}^{2-} \rightarrow \text{V}_2\text{S}_5$

17) chromium (III) hydroxide  $\text{Cr}^{+3}(\text{OH})_3$

18) lithium iodide  $\text{Li}^+\text{I}^-$

19) lead (II) nitride  $\text{Pb}^{+2}\text{N}^{3-} \rightarrow \text{Pb}_3\text{N}_2$

20) silver bromide  $\text{AgBr}$